

CHEMISTRY STRUCTURE PROPERTIES NIVALDO TRO



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In chemistry and atomic physics, the electron affinity (E_{ea}) of an atom or molecule is defined as the amount of energy released or spent when an electron is added to a neutral atom or molecule in the gaseous state to form a negative ion.. $X + e \rightarrow X^- + \text{energy}$. In solid state physics, the electron affinity for a surface is defined somewhat differently (see below).

Electron affinity - Wikipedia

A hydrogen bond is the attraction between the lone pair of an electronegative atom and a hydrogen atom that is bonded to either nitrogen, oxygen, or fluorine. The hydrogen bond is often described as a strong electrostatic dipole–dipole interaction. However, it also has some features of covalent bonding: it is directional, stronger than a van der Waals force interaction, produces interatomic ...

Intermolecular force - Wikipedia

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